Past Paper Questions

Class 9 Chemistry · Chapter 2 — Matter

Paper	Source
Model Paper 2025	FBISE Model Paper · Curriculum 2022-23 · Inclusive Scheme of Studies 2024
2025, 1st Exam Paper 1	FBISE SSC-I · 2025 · Versions B and D · 1st Exam
2025, 1st Exam Paper 2	FBISE SSC-I · 2025 · Version D · 1st Exam (incl. Old Curriculum 2006 pages)
2025, 2nd Exam	FBISE SSC-I · 2025 · 2nd Exam
Model Paper 2026 (50% MCQs)	FBISE Model Paper · National Curriculum 2022-23 · 50% MCQs + 50% Descriptive
2026, Paper D	FBISE SSC-I · 2026 · Paper D
2026, Paper M	FBISE SSC-I · 2026 · Paper M
2026, Paper B	FBISE SSC-I · 2026 · Paper B

Chapter 2 questions taken from Federal Board SSC-I Chemistry papers (2025 and 2026), arranged paper by paper. Marks shown in teal brackets. “OR” marks the alternative question in a choice pair. Questions are given without answers. Total: 38 questions.

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Model Paper 2025

Curriculum 2022-23, Inclusive Scheme of Studies 2024

Section A — MCQ

Q1(xii). Suggest which statement about solubility is INCORRECT?

A) It refers to the ability of a solute to dissolve in a solvent
B) It is measured in grams per liter of solvent
C) It is affected by temperature, pressure, and the nature of the solute and solvent.
D) It is independent of temperature.

Section B — Short Questions

Q2(iii). Compare following properties of liquid Crystals with liquids.

a. Molecular arrangement
b. Fluidity **[1.5+1.5]**

Section C — Long Questions

Q5. Define solubility. How can it be affected by the change in temperature? Explain with graph. **[1+2+3]**

2025, 1st Exam Paper 1

Versions B and D

Section A — MCQ

Q12 (Versions B and D). In which state of matter do particles have the least compressibility?

A) Solid B) Liquid C) Gas D) Plasma

Section B — Short Questions

Q2(ii) OR (Version B). Briefly describe the following properties of gas:

a) Density
b) Fluidity **[1.5+1.5]**

Section C — Long Questions

Q6 (Version B). What is meant by allotropy? Discuss any two allotropes of carbon. **[1+1.5+1.5]**

2025, 1st Exam Paper 2

Version D, Sections B and C

Section B — Short Questions

Q2(iii). Briefly describe the following properties of solids:

a) Compressibility
b) Density **[3]**

Section C — Long Questions

Q6. Compare colloid and suspension. Give one example of each. **[2+2]**

Old Curriculum 2006 pages

These pages follow the older syllabus. Only questions matching Chapter 2 topics are included.

Q4 (MCQ). Which of the following list contains only elements?

A) Water, Oxygen
B) Hydrogen, Oxygen, Salt
C) Air, Water, Fire
D) Sulphur, Carbon

Q8 (MCQ). Amalgam used by dentist is an example of:

A) Solid in liquid
B) Solid in solid
C) Liquid in solid
D) Liquid in liquid

Q11 (MCQ). 10g sugar in 100g water is an example of:

A) Solvent B) Colloid C) Suspension D) Aqueous solution

Q2(iv) OR (Short). Predict the solubility of following substances:

a. Methanol in water
b. Oil in benzene
c. Sodium chloride in water **[3]**

Q2(viii) (Short). Describe colloids and suspension with one suitable example of each. **[3]**

Q2(viii) OR (Short). What is meant by solution and percentage of solution? **[3]**

Q2(ix) (Short). Briefly explain any two allotropic forms of carbon. **[3]**

2025, 2nd Exam

Section A — MCQ

Q12. Most compressible state of matter is:

A) Gas B) Liquid C) Solid D) Plasma

Section B — Short Questions

Q2(iii). What is meant by BEC? How are BECs formed? **[1+2]**

Section C — Long Questions

Q6. Compare solutions and colloids. (Any three points) Give one example of each. **[3+1]**

Model Paper 2026 (50% MCQs)

National Curriculum 2022-23, 50% MCQs and 50% Descriptive

Section A — MCQ

Q1. Diamond is used in cutting tools. The reason for this use is:

A) It has a layered structure with weak van der Waals forces
B) It is a highly dense form of pure carbon
C) It has a strong three-dimensional network structure
D) It contains delocalized free electrons

Q2. Which of the following is a compound?

A) Air B) Iron C) Carbon dioxide D) Brass

Q3. Which of the following combinations is correctly matched?

A) Sugar + Water → Suspension
B) Milk → Colloid
C) Sand + Water → Solution
D) Air → Suspension

Scenario-Based MCQs (Q45–Q48)

> **Scenario:** Ali melts a piece of wax in a test tube by heating it in a water bath. After removing the heat source, he observes that the liquid wax slowly turns back into a solid.

Q45. What is the name of the process when liquid wax turns into a solid?

A) Condensation B) Evaporation C) Freezing D) Melting

Q46. Why is this change considered a physical change?

A) A new substance is formed
B) The chemical composition changes
C) The change is reversible and no new substance is formed
D) Heat is released permanently

Q47. During this change, what happens to the particles of wax?

A) They break into new atoms
B) They become chemically different
C) They come closer together without changing identity
D) They disappear

Q48. Which property remains unchanged during this process?

A) State B) Shape C) Chemical composition D) Volume

Section B — Short Questions

Q1 OR. State the difference between a compound and a mixture. In your answer, briefly describe the composition of each type of substance. **[3]**

Q6 OR. How different states of matter (solids, liquids, gases, and plasma) are involved in chemical interactions. **[2]**

The "Define Chemistry" part belongs to Chapter 1.

2026, Paper D

Section A — MCQ

Q1. Which state of matter is represented by air?

A) Solid B) Liquid C) Gas D) Plasma

Q2. Diamond and graphite are allotropic forms of:

A) Oxygen B) Carbon C) Silicon D) Sulfur

Section B — Short Questions

Q2(i) OR. Differentiate between suspension and colloidal with one suitable example each. **[1.5+1.5]**

Section C — Long Questions

Q3. Differentiate between compounds, and mixtures with respect to:

a. Composition
b. Chemical bonding
c. Separation methods **[2+2+2]**

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2026, Paper M

Section B — Short Questions

Q2(i) OR. How is solubility affected by increasing temperature for given solutions:

a. KCl in water (endothermic solution)

b. $Ca(OH)_2$ in water (exothermic solution) **[1.5+1.5]**

Q2(x) OR. Briefly describe the compressibility of solids, liquids and gases with reason. **[1+1+1]**

Section C — Long Questions

Q3. Compare diamond and graphite in terms of their structure and properties. **[3+3]**

2026, Paper B

Section B — Short Questions

Q2(ii). Why solids have high density and very low compressibility whereas gases have opposite behavior? Explain briefly. **[1.5+1.5]**

Q2(iv). Identify the given substances as solution, suspension or colloidal:

a. Air

b. Sand in water

c. Jelly **[1+1+1]**

Q2(v). What is meant by 'element'? State one distinguishing feature that separates it from mixtures. **[2+1]**

Q2(vii). How is solubility affected by increasing temperature for given solutions:

a. KCl in water (endothermic solution)

b. $Ca(OH)_2$ in water (exothermic solution) **[1.5+1.5]**

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